

Chem 6
Homework 6 -- Intermolecular Forces

Zumdahl 5th ed. Ch. 16, exercises 11, 15-23 (odd), 29, 33, 91, 93, 97, and Ch. 13, 25-29 (odd)

1. Which of the following would you expect to show London dispersion forces? dipole-dipole interactions? hydrogen bonding? (a) CH₄ (b) CH₃Cl (c) CH₂Cl₂ (d) CHCl₃ (e) CCl₄
2. Which of the following would you expect to show dipole-dipole interactions?
(a) H₂S (b) CO₂ (c) CF₄ (d) CO
3. On the basis of their intermolecular interactions, arrange the following substances in order of increased boiling points: NaF, Al₂O₃, Ar, NO, CH₄